I. Fill in the blanks with the most appropriate term:

A	tells the story of a chemical
reaction	are the starting substances in the reaction
while	are the new substances that are formed. The
large numbers in	front of some of the formulas are called
·	These numbers are used to the
equation because	chemical reactions must obey the Law of
	of Matter. The number of atoms of each element on
both sides of the eq	uation must be because matter
cannot be	or When balancing
equations, the only numbers that can be changed are;	
remember that	must never be changed in order to
balance an equation.	

II. Balance the following equations:

- 1. Al + $O_2 \rightarrow Al_2O_3$
- 2. $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$
- 3. $AI(NO_3)_3 + NaOH \rightarrow AI(OH)_3 + NaNO_3$

4.
$$KNO_3 \rightarrow KNO_2 + O_2$$

5.
$$O_2$$
 + CS_2 \rightarrow CO_2 + SO_2

6.
$$KCIO_3 \rightarrow KCI + O_2$$

7.
$$BaF_2 + K_3PO_4 \rightarrow Ba_3(PO_4)_2 + KF$$

8.
$$H_2SO_4 + Mg(NO_3)_2 \rightarrow MgSO_4 + HNO_3$$

9. Al +
$$H_2SO_4 \rightarrow Al_2(SO_4)_3 + H_2$$

10.
$$WO_3 + H_2 \rightarrow W + H_2O$$